



ST. LAWRENCE HIGH SCHOOL
A JESUIT CHRISTIAN MINORITY INSTITUTION
SOLUTION-35(CLASS-12)
TOPIC- ELECTROCHEMISTRY
SUBTOPIC- FARADAY'S LAW



SUBJECT – CHEMISTRY
DURATION – 30 mins

F.M. - 15
DATE -02.07.20

1.1 Amount of electricity that can deposit 108g silver from AgNO₃ solution is:

- (a) 1 ampere (b) 1 coulomb (c) 1 faraday (d) None of the above

Ans. c

1.2 When 9.65 Coulomb of electricity is passed through a solution of silver nitrate, the amount of silver deposited is:

- (a) 10.8 mg (b) 5.4 mg (c) 16.2 mg (d) 21.2 mg

Ans. a

1.3 Three Faradays electricity was passed through an aqueous solution of Iron (II) bromide. The weight of Iron metal, deposited at the cathode (in gram):

- (a) 56 (b) 84 (c) 112 (d) 168

Ans. b

1.4 A silver cup is plated with silver by passing 965 coulombs of electricity, the amount of silver deposited is:

- (a) 9.89 g (b) 107.87 g (c) 1.0787 g (d) 1.002g

Ans. c

1.5 The atomic mass of Al is 27u, When a current of 5 Faradays is passed through a solution of Al³⁺, the mass of Al deposited is:

- a) 27g b) 36g c) 45g d) 39g

Ans. c

1.6 An apparatus used for the measurement of quantity of electricity is known as:

- (a) Calorimeter (b) Cathetometers (c) Coulometer (d) Colorimeter

Ans. c

1.7 The unit of electrochemical equivalent is:

- (a) Gram (b) Gram/ampere (c) Gram/coulomb (d) Coulomb / Gram

Ans. c

1.7 On electrolysis of dilute sulphuric acid using platinum electrodes, the product obtained at the anode will be-

a) Hydrogen b) oxygen c) hydrogen sulphide d) Sulphur dioxide

Ans. b

1.8 A certain amount of current liberates 0.504g of Hydrogen in 2 hours. How many gram of copper can be liberated by the same time in a copper sulphate solution?

(a) 12.7g (b) 15.9g (c) 31.8g (d) 63.5g

Ans. b

1.9 A device that converts energy of combustion of fields like hydrogen and methane directly into electrical energy is known as-

a) Electrolytic cell b) Dynamo c) Ni-Cd cell d) Fuel cell

Ans. d

1.10 The reciprocal of electrical resistance is-

a) Voltage b) current c) conductance d) none of the above

Ans. c

1.11 The unit of faraday is:

a) Ampere b) Coulomb c) Coulomb/mol d) Coulomb/sec

Ans. c

1.12 Without losing its concentration $ZnCl_2$ solution cannot be kept in contact with-

a) Au b) Al c) Pb d) Ag

Ans. b

1.13 The factors which determine the flow of electricity through a solution-

a) Nature of electrolyte or interionic attractions b) The nature of the solvent and its viscosity c) Potential difference d) Both a and b

Ans. d

1.14 As temperature increases electrolytic conduction-

a) Increases b) Decreases c) Remains unaffected d) None of the above

Ans. a

1.15 The desired amount of charge for obtaining 1 mole of Al from Al^{3+} :

a) 3×96500 C b) 96500 C c) $96500/3$ C d) $96500/2$ C

Ans. a

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