ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION

WORKSHEET-08(CLASS-11)

TOPIC- SOME BASIC CONCEPT OF CHEMISTRY

SUBTOPIC- CALCULATION OF STRENGTH

SUBJECT – CHEMISTRY
DURATION – 30 mins

F.M. - 15
DATE -23.06.20

1.1 **1 M of $Fe(OH)_3$ Solution is -**

a) 2N b) 3N c) 1N d) 0.333N

1.2 **0.5 gram equivalent of H_2S is equal to-**

a) 0.25 Moles of H_2S b) 1 mole of H_2S c) 0.05 moles of H_2S d) None of the above

1.3 **300 ml 0.2 M HCl and 200 ml of 0.03M H_2SO_4 are mixed. The normality of the resulting mixture will be-**

a) 0.044 N b) 0.72 N c) 0.84 N d) 0.144 N

1.4 **Find the Normality of the solution containing 0.5 gm of NaOH in 1L solution-**

a) 0.0125 N b) 0.125 N c) 0.5 N d) 0.0250 N

1.5 **How much water is to be added to prepare a 0.25N HCL solution from 0.5 N HCL 1 L solution-**

a) 500 ml b) 100 ml c) 1000 ml d) 250 ml

1.6 **20 ml of a 0.125 N HCl solution were neutralised by 25 ml of a KOH solution. What is the normality of KOH solution?**

a) 0.01N b) 0.1 N b) 0.2 N d) None of the above

1.7 **Calculate the normality of 0.321 g sodium carbonate when it is mixed in a 250 mL solution.**

a) **0.0755 N** b) **0.0725 N** c) **0.0355 N** d) **0.0955 N**

1.8 **Calculate the number of moles & molarity of N/2 500 ml solution of H_2SO_4-**

a) 0.125 and 0.25M b) 0.25 and 0.125M c) 0.25 and 0.25M d) 0.125 and 0.125M

1.9 **Calculate the normality of NaOH solution Formed by dissolving 0.2 gm NaOH to make 250 ml solution-**

a) 0.3N b) 0.2N c) 0.4N d) 0.35N

1.10 **Which among the following is temperature independent?**

a) Normality b) Molarity c) Formality d) Molality

1.11 **What is the volume of a 0.2 M $AgNO_3$ solution containing 8.5 grams of $AgNO_3$?**

a) 0.25 L or 150 ml b) 0.25 L or 250 ml c) 0.15 L or 150 ml d) 0.75 L or 750 ml

1.12 **How many grams of $KMnO_4$ is required to prepare 1.0 L of a solution of 1.5 M $KMnO_4$?**

a) 477g b) 230g c) 237 d) 327g

1.13 **What is the molarity of a solution containing 72.06 grams of $BaCl_2$ in enough water to make 800 mL of solution?**

a) 0.133 M b) 0.336 M c) 0.433 M d) 0.633 M

1.14 **What is the molarity of a solution containing 11.522 grams of KOH in enough water to make 350 mL of solution?**

a) 0.586 M b) 0.786 M c) 0.625 M d) 0.725 M

1.15 **What will the concentration of citric acid be if 25.00 ml of the citric acid solution is titrated with 28.12 mL of 0.1718 N KOH?**

a) 0.2126 N b) 0.2332 N c) 0.1236 N d) 0.1932 N

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