



# ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION



**Sub: Physical Science**

**Class: 8**

**Date: 30.06.20**

**Duration: 40 min**

**Worksheet 51 Full Marks: 15**

## LANGUAGE OF CHEMISTRY

**Choose the Correct options:**

1. Calcium ion ( $\text{Ca}^{2+}$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
2. Sodium ion ( $\text{Na}^+$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
3. Potassium ion ( $\text{K}^+$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
4. Cuprous ion ( $\text{Cu}^+$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
5. Ferric ion ( $\text{Fe}^{3+}$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
6. Chloride ion ( $\text{Cl}^-$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
7. Sulphide ion ( $\text{S}^{2-}$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
8. A negatively charged atom is called
  - a. Anion
  - b. Cation
  - c. Radical
9. Valency is the combining capacity of
  - a. Atoms
  - b. Molecules
  - c. Electrons
10. Hydrogen ion ( $\text{H}^+$ ) is
  - a. Monovalent
  - b. Divalent
  - c. Trivalent
11. Hydrogen ion ( $\text{H}^+$ ) is in
  - a. Octet state
  - b. Duplet state
  - c. Unstable

12. Hydroxyl ion ( $\text{OH}^-$ ) is a
- Radicle
  - Radical
  - Colloid
13. The L shell contains
- 2 electrons
  - 8 electrons
  - 18 electrons
14. Octet rule states that
- The last shell of a stable atomic configuration has eight electrons
  - If the first shell is the last shell of an atomic configuration it has two electrons
  - Both
15. The L shell contains 8 electrons according to
- Octet Rule
  - Duplet Rule
  - Bohr-Bury Rule