



## ST. LAWRENCE HIGH SCHOOL



A JESUIT CHRISTIAN MINORITY INSTITUTION

• **Subject Physical science Worksheet- 6 Class 7**

• **Date 16.05.2020**

• **Chapter: Atoms, Molecules and Radicals**

• **Answer the following questions (MCQ) : (1×15)**

• 1

Q. How many valence electrons are found in atoms of group 14?

answer choices

4

3

14

16

• Question 2

Q. Which group has the greatest number of valence electrons?

answer choices

1

14

18

16

• Question 3

Q. How many valence electrons are found in atoms of group 15?

answer choices

4

5

2

3

• Question 4

Q. How many valence electrons are found in atoms of group 1?

answer choices

7

6

4

1

• Question 5

Q.

What are valence electrons?

[answer choices](#)

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The total number of electrons in an atom

The number of electrons in the outermost shell

The number of electrons in the second shell

The number of protons in the outermost shell

• Question 6

Q.

Which elements have the most similar chemical properties?

[answer choices](#)

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boron and carbon

oxygen and sulfur

aluminum and bromine

argon and silicon

• Question 7

Q.

All atoms are most stable with (or would "prefer") how many electrons in their valence shell?

[answer choices](#)

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1

2

8

18

• Question 8

Q. Valence electrons are:

[answer choices](#)

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Electrons farthest away from the nucleus

Electrons closest to the nucleus

Electrons that just come and go - they don't stay with the atom

Electrons in the second shell

• Question 9

Q.

When trying to identify an unknown element, a scientist determines what other elements reacts with chemically. Which property of the unknown element determines the other element it reacts with?

[answer choices](#)

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The total number of neutrons in the unknown element

The total number of particles in the nucleus of the unknown element

The number of protons in the nucleus of the unknown element

The number of valence electrons in the unknown element

• Question 10

Q.

What groups are the most reactive metals?

[answer choices](#)

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2 and 3

1 and 18

1 and 2

16 and 17

• Question 11

Q.

What groups are the most reactive non metals?

[answer choices](#)

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1 and 2

16 and 17

2 and 17

1 and 16

• Question 12

Q.

Why are halogens so reactive?

[answer choices](#)

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They want to get rid of their only valance electron

They only need one more electron

They are non reactive, they have a full shell

- Question 13

Q.

How do electrons of an element determine that element's reactivity?

[answer choices](#)

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If the valence shell has missing electrons it is very reactive to attract or lose more electrons.

If the valence shell has a full ring of electrons it is very reactive because it needs more.

If the valence shell has a full shell of electrons it is very reactive because it is already full.

If the valence shell has missing electrons it is not reactive because it is missing electrons.

- Question 14

Q. Which of the following statements describes group 2 of the periodic table?

[answer choices](#)

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Stable non-metals with 1 valence electron

Stable Metals with 2 valence electrons

Highly reactive non-metals with 1 valence electron.

Highly reactive metals with 2 valence electrons.

- Question 15

Q. Describe this element.

[answer choices](#)

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Highly reactive metal

Highly reactive nonmetal

Stable

Less stable non metal

Teacher- PiyaliHalder