

## ST. LAWRENCE HIGH SCHOOL A JESUIT CHRISTIAN MINORITY INSTITUTION WORKSHEET-35(CLASS-12) TOPIC- ELECTROCHEMISTRY SUBTOPIC- FARADAY'S LAW



SUBJECT – CHEMISTRY DURATION – 30 mins F.M. - 15 DATE -02.07.20

1.1 Amount of electricity that can deposit 108g silver from AgNO3 solution is:

(a) 1 ampere (b) 1 coulomb (c) 1 faraday d) None of the above (c)

1.2 When 9.65 Coulomb of electricity is passed through a solution of silver nitrate, the amount of silver deposited is: (a)
(a) 10.8 mg (b) 5.4 mg (c) 16.2 mg (d) 21.2 mg

1.3 Three Faradays electricity was passed through an aqueous solution of Iron (II) bromide.The weight of Iron metal, deposited at the cathode (in gram): (b)(a) 56 (b) 84 (c) 112 (d) 168

1.4 A silver cup is plated with silver by passing 965 coulombs of electricity, the amount of silver deposited is: (c) (a) 9.89 g (b) 107.87 g (c) 1.0787 g (d) 1.002g

1.5 The atomic mass of Al is 27u, When a current of 5 Faradays is passed through a solution of Al<sup>3+</sup>, the mass of Al deposited is: © a) 27g b) 36g c) 45g d) 39g

1.6 An apparatus used for the measurement of quantity of electricity is known as: © (a) Calorimeter (b) Cathetometers (c) Coulometer (d) Colorimeter

1.7 The unit of electrochemical equivalent is: ©(a) Gram (b) Gram/ampere (c) Gram/coulomb (d) Coulomb / Gram

1.7 On electrolysis of dilute sulphuric acid using platinum electrodes, the product obtained at the anode will be-

a) Hydrogen b) oxygen c) hydrogen sulphide d) Sulphur dioxide

1.8 A certain amount of current liberates 0.504g of Hydrogen in 2 hours. How many gram of copper can be liberated by the same time in a copper sulphate solution? (b) (a) 12.7g (b) 15.9g (c) 31.8g (d) 63.5g

# 1.9 A device that converts energy of combustion of fields like hydrogen and methane directly into electrical energy is known as-

a) Electrolytic cell b) Dynamo c) Ni-Cd cell d) Fuel cell

#### 1.10 The reciprocal of electrical resistance is-

a) Voltage b) current c) conductance d) none of the above

#### 1.11 The unit of faraday is: ©

a) Ampere b) Coulomb c) Coulomb/mol d) Coulomb/sec

#### 1.12 Without losing its concentration ZnCl<sub>2</sub> solution cannot be kept in contact with-

a) Au b) Al c) Pb d) Ag

#### 1.13 The factors which determine the flow of electricity through a solution-

a) Nature of electrolyte or interionic attractions b) The nature of the solvent and its viscosity c) Potential difference d) Both a and b

#### 1.14 As temperature increases electrolytic conduction-

a) Increases b) Decreases c) Remains unaffected d) None of the above

# 1.15 The desired amount of charge for obtaining 1 mole of Al from $Al^{3+}$ : (a)

a) 3X96500 C b) 96500 C c) 96500/3 C d) 96500/2 C

## PREPARED BY: MR. ARNAB PAUL CHOWDHURY