



# ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION

## WORKSHEET-30(CLASS-11)

### TOPIC- PERIODIC PROPERTIES

#### SUBTOPIC-BASIC CONCEPT

**SUBJECT – CHEMISTRY**

**DURATION – 30 mins**

**F.M. - 15**

**DATE -08.08.20**



**1. 14 elements after actinium is called-**

- a. Lanthanides
- b. Actinides
- c. d-block elements
- d. P-block elements

**2. An element has an atomic number of 15 with which of the following elements will it show similar chemical properties-**

- a. Be (4)
- b. Ne (10)
- c. N(7)
- d. O (8)

**3. The group number and period number respectively of an element with atomic number 8 is-**

- a. 6,2
- b. 16,2
- c. 6,8
- d. 16,4

**4. An element belongs to period 2 and group 2 th number of valence electrons in the atoms of this element is-**

- a. 2
- b. 4
- c. 3
- d. 1

**5. In the third period of the periodic table the element having smallest size is-**

- a. Na
- b. Ar
- c. Cl
- d. Si

**6. Electronic configuration of  $\text{Al}^{+3}$  is-**

- a. 2,8,3
- b. 2,8,8
- c. 2,8
- d. 2,8,8,3

**7. Identify the group which is not a Dobereiner triad-**

- a. Li, Na, K
- b. Be, Mg, Cr
- c. Ca, Sr, Ba
- d. Cl, Br, I

**8. Which is not true about the noble gases?**

- a. They are non-metallic in nature
- b. They exist in atomic form
- c. They are radioactive in nature
- d. Xenon is the most reactive among these

**9. Identify the wrong sequence of the elements in a group-**

- a. Ca, Br, Ba
- b. Cu, Au, Ag
- c. N, P, As
- d. Cl, Br, I

**10. An element with atomic number will form a basic oxide\_\_\_\_\_**

- a. 7
- b. 17
- c. 14
- d. 11

**11. Which among the following is not an inherent property of an element?**

- a. Ionisation energy
- b. Electron affinity
- c. Electronegativity
- d. Valency

**12. According to Lothar Meyer, the \_\_\_\_\_ is the periodic function of physical properties of an element-**

- a. Ionisation energy
- b. Atomic number
- c. Atomic weight
- d. Atomic radius

**13. Which among the following is not a periodic property of an element?**

- a. Ionisation energy
- b. Electron affinity
- c. Electronegativity
- d. Radioactivity

**14. Which among the following set of elements can be considered as coinage metals?**

- a. Cu, Ag, Au
- b. W, Cu, Pt
- c. Cu, Fe, Au
- d. Na, Cu, Hg

15. Which among the following has the least electronegativity at Pauling scale among the following?

- a. F
- b. Cs
- c. Mg
- d. N

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