



ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION



Sub: Physical Science

Class: 8

Date: 28.11.20

Duration: 40 min

Worksheet Solution 07

Full Marks: 15

LANGUAGE OF CHEMISTRY: TYPES OF CHEMICAL REACTIONS

Choose the Correct options:

- When two or more reactants combine to form a product
 - Combination reaction**
 - Displacement reaction
 - Double displacement reaction
 - Decomposition reaction
- When one substance breaks into two or more substances
 - Combination reaction
 - Displacement reaction
 - Double displacement reaction
 - Decomposition reaction**
- One displacement displaces another from its compound
 - Combination reaction
 - Displacement reaction**
 - Double displacement reaction
 - Decomposition reaction
- Positive and negative radical of reactants are exchanged
 - Combination reaction
 - Displacement reaction
 - Double displacement reaction**
 - Decomposition reaction
- When one substance breaks into two or more substances by passing electric current
 - Electrolysis**
 - Rusting
 - Neutralization
 - None of these
- Formation of crust on a piece of iron in the presence of moist air
 - Electrolysis
 - Rusting**
 - Neutralization
 - None of these
- An acid reacts with a base to form salt and water
 - Electrolysis
 - Rusting
 - Neutralization**
 - None of these
- Which of the following can displace hydrogen from a dilute acid?
 - Silver
 - Copper
 - Zinc**
 - Mercury
- What are reactants?
 - chemicals start the reaction**
 - chemicals the reaction produced
 - chemicals on the right side of the arrow
 - all of the above
- What are the products? $\text{Cu} + \text{AgNO}_3 \rightarrow$
 - $\text{Cu}(\text{NO}_3)_2 + \text{Ag}$**
 - $\text{Cu}(\text{NO}_3) + \text{Ag}$

- c. $\text{CuAg} + \text{NO}_3$
d. $\text{Ag} + \text{CuNO}_2$
11. What kind of reaction is this?: $\text{SO}_2 + \text{O}_2 \rightarrow \text{SO}_3$
a. **Combination reaction**
b. Displacement reaction
c. Double displacement reaction
d. Decomposition reaction
12. What kind of reaction is this?: $\text{H}_2\text{CO}_3 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
a. Combination reaction
b. Displacement reaction
c. Double displacement reaction
d. **Decomposition reaction**
13. What kind of reaction is this?: $\text{H}_2\text{SO}_4 + \text{Ca} \rightarrow \text{CaSO}_4 + \text{H}_2$
a. Combination reaction
b. **Displacement reaction**
c. Double displacement reaction
d. Decomposition reaction
14. What kind of reaction is this?: $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
a. Combination reaction
b. Displacement reaction
c. **Double displacement reaction**
d. Decomposition reaction
15. What kind of reaction is this? $\text{BaO} + \text{HNO}_3 \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{H}_2\text{O}$
a. Combination reaction
b. Displacement reaction
c. **Double displacement reaction**
d. Decomposition reaction