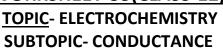


## ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION

## WORKSHEET-33(CLASS-12)





SUBJECT – CHEMISTRY DURATION – 30 mins

F.M. - 15 DATE -30.06.20

- 1.1 Aqueous solution of which of the following compounds is the best conductor of electric current?
- (a) Acetic acid (b) Hydrochloric acid (c) Ammonia d) Fructose
- 1.2 The unit of conductance-
- (a) Siemens (b) mho (c) Ohm (d) Both a and b
- 1.3 An increase in the conductivity equivalent of a solid electrolyte with dilution is primarily due to-
- (a) Increased ionic mobility of ions (b) 100 percent electrolyte ionisation with natural dilution (c) increase in both ion numbers and ionic mobility (d) A rise in ion counts
- 1.4 The ionic conductance of  $Ba^{2+}$  and  $Cl^-$  are respectively 127 and 76 ohm<sup>-1</sup> at infinite dilution. The equivalent conductance of  $BaCl_2$  at infinite dilution will be-
- (a) 139 (b) 203 (c) 279 (d) None of these
- 1.5 When heating one end of a metal plate, the other end gets hot because of -
- (a) Resistance of the metal (b) mobility of atoms in the metal (c) energised electrons moving to the other end (d) minor perturbation in the energy of atoms.
- 1.6 The weight of silver displaced by a quantity of electricity which displaces 5600ml of O<sub>2</sub> at STP will be-
- (a) 5.4g (b) 10.8g (c) 54.9g (d) 108.0g
- 1.7 On electrolysis of dilute sulphuric acid using platinum electrodes, the product obtained at the anode will be-
- a) Hydrogen b) oxygen c) hydrogen sulphide d) Sulphur dioxide
- 1.8 Once a current of 1.0 ampere was passed through one liter of  $CuCl_2$  solution for 16 min and 5sec, all of the solution's copper was deposited at cathode. The power of solution  $CuCl_2$  was (Molar mass of Cu = 63.5; Faraday constant = 96,500 C/ mol)-
- (a) 0.01N (b) 0.01M (c) 0.02M (d) 0.2N

- 1.9 A device that converts energy of combustion of fields like hydrogen and methane directly into electrical energy is known as-
- a) Electrolytic cell b) Dynamo c) Ni-Cd cell d) Fuel cell
- 1.10 Which reaction is not feasible?
- a)  $2KO + Br_2 \rightarrow 2KBr + I_2$  b)  $2KBr + I_2 \rightarrow 2KI + Br_2$  c)  $2KBr + CI_2 \rightarrow 2KCI + Br_2$
- d)  $2H_2O + 2F_2 \rightarrow 4HF + O_2$
- 1.11 The unit of equivalent conductivity-
- a) S cm<sup>2</sup> equivalent<sup>-1</sup> b) S cm<sup>2</sup> equivalent c) S cm<sup>3</sup> equivalent d) S cm equivalent
- 1.12 Without losing its concentration ZnCl<sub>2</sub> solution cannot be kept in contact with-
- a) Au b) Al c) Pb d) Ag
- 1.13 The factors which determine the flow of electricity through a solution-
- a) Nature of electrolyte or interionic attractions b) The nature of the solvent and its viscosity c) Potential difference d) Both a and b
- 1.14 Electrode potential of a cell is-
- a) An intensive property b) Extensive property c) Both a and b d) can't be predicted
- 1.15 The conductivity of electrolytic (ionic) solutions depends on-
- a) the nature of the electrolyte added b) temperature c) size of the ions produced and their solvation d) all of these

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