



## ST. LAWRENCE HIGH SCHOOL



A JESUIT CHRISTIAN MINORITY INSTITUTION

• **Subject Physical science** \_\_\_\_\_ **Worksheet- 7** **Class 7**

• **Date 18.05.2020**

• **Chapter :Atoms,Molecules and Radicals**

• **Answer the following questions (MCQ) :**

**(1×15)**

• Question1

Q. What is the empirical formula of the following molecular formula:  $C_3H_6$

answer choices

$C_3H_6$

$CH_2$

$C_2H_4$

$CH_3$

Question2

Q. What is the molecular formula if the empirical formula is  $C_2H_5$  and the molecular molar mass is 58.14 g/mol?

answer choices

$C_2H_5$

$C_4H_{10}$

$C_1H_{2.5}$

$C_4H_8$

Question 3

Q. What is the empirical formula for the following molecular formula:  $C_6H_{14}$

answer choices

$C_6H_{14}$

$C_3H_7$

$CH_2$

$CH_3$

• Question 4

Q. What is the molecular formula if the empirical formula is  $CH_2O$  and the molecular molar mass is 180.18?

answer choices

$CH_2O$

$C_2H_4O_2$

$C_4H_8O_4$

$C_6H_{12}O_6$

Question 5

Q. What is the empirical formula for the following molecular formula:  $C_5H_{12}$

answer choices

$C_5H_{12}$

$CH_3$

$CH_2$

$C_{2.5}H_6$

Question 6

Q. Which pair has the same empirical formula?

answer choices

NaCrO<sub>4</sub> and Na<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>  
C<sub>2</sub>H<sub>4</sub>O<sub>2</sub> and C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>  
C<sub>3</sub>H<sub>6</sub>O<sub>3</sub> and C<sub>2</sub>H<sub>6</sub>O<sub>2</sub>  
CH<sub>4</sub> and C<sub>2</sub>H<sub>6</sub>

Question 7

Q. A 4.50g sample of which of the following would have the greatest mass of oxygen?

answer choices

Na<sub>2</sub>O (molar mass = 62g/mol)

Li<sub>2</sub>O (molar mass = 30g/mol)

MgO (molar mass = 40g/mol)

SrO (108g/mol)

Question 8

Q. A compound is determined to contain 14g of nitrogen and 32g of oxygen. The empirical formula of the compound is:

answer choices

NO

NO<sub>2</sub>

N<sub>2</sub>O

NO<sub>3</sub>

- Question 9

Q. How many moles of carbon are in 88g of propane, C<sub>3</sub>H<sub>8</sub>.

answer choices

2.0

16.0

6.0

96.0

- Question 10

Q. 2H<sub>2</sub>O

How many H<sub>2</sub>O molecules are present?

answer choices

1

4

2

8

- Question 11

Q. What do we call a chemical substance made up of **2 or more different types atoms** bonded together?

answer choices

formula

atom

subscript

compound.

Question 1 2

Q. All \_\_\_\_\_ are made up of two or more atoms bonded together.

answer choices

atoms

molecules

electrons  
elements

Question 13

Q. Bonds between atoms in a molecule form as a result of the sharing of?

answer choices

formulas  
electrons  
atoms  
chemicals

Question 14

Q. A recipe for a chemical substance is called a \_\_\_\_\_.

answer choices

coefficient  
chemical formula  
Oxygen  
Atom

Question 15

Q.  $H_2O_2$  ← ?

This number tells the number of atoms present.

answer choices

coefficient  
molecule  
subscript  
atom

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