



# ST. LAWRENCE HIGH SCHOOL

A JESUIT CHRISTIAN MINORITY INSTITUTION

## WORKSHEET-32(CLASS-11)

### TOPIC- PERIODIC PROPERTIES

#### SUBTOPIC-BASIC CONCEPT



**SUBJECT – CHEMISTRY**

**DURATION – 30 mins**

**F.M. - 15**

**DATE -15.08.20**

**1. Find the set of triad-**

- (a) Li, Be, B
- (b) H, He, B
- (c) B, C, O
- (d) He, Li, B

**2. In the modern periodic table, the number of group of the element is the same as-**

- (a) principal quantum number
- (b) atomic number
- (c) azimuthal quantum number
- (d) none of these

**3. The correct order for the size of I, Br and F-**

- (a)  $I > Br > F$
- (b)  $Br > F > I$
- (c)  $Br > I > F$
- (d)  $Br > I < F$

**4. For the same value of n, the screening power of orbital follows the order**

- (a)  $s = p = d = f$
- (b)  $p > s > d > f$
- (c)  $f < d < p < s$
- (d)  $s < p < d < f$

**5. Which of the reactions will need the minimum amount of energy?**

- (a)  $Na \rightarrow Na^+ + e^-$
- (b)  $Ca^+ \rightarrow Ca^{++} + e^-$
- (c)  $K^+ \rightarrow K^{++} + e^-$
- (d)  $C^{2+} \rightarrow C^{3+} + e^-$

**6. Which of the following statements is incorrect?**

- (a) I.E.<sub>1</sub> of O is lower than that of N but I.E.<sub>2</sub> O is higher than that of N
- (b) The enthalpy of N to gain an electron is almost zero but of P is 74.3 kJ mol<sup>-1</sup>
- (c) isoelectronic ions belong to the same period
- (d) The covalent radius of iodine is less than its Van der Waal's radius

**7. The correct order of electron affinity is-**

- (a) Cl > F > Br > I
- (b) F > O > Cl > Br
- (c) F > Cl > Br > O
- (d) O > F > Cl > Br

**8. The pair which doesn't obey old formed Mendeleev periodic table-**

- (a) Ar & K
- (b) Ar & Cl
- (c) Na & K
- (d) C & N

**9. Which one is the most basic among these?**

- (a) MgO
- (b) CaO
- (c) Al<sub>2</sub>O<sub>3</sub>
- (d) Na<sub>2</sub>O

**10. Which one will have the highest 1st ionisation energy?**

- (a) 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>1</sup>
- (b) 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>4</sup>
- (c) 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup>
- (d) 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup>

**11. Elements in the same vertical group of the periodic table have same-**

- (a) Number of valence electrons
- (b) Atomic number
- (c) Atomic mass
- (d) Atomic volume

**12. Which period contains minimum number of elements?**

- (a) 5
- (b) 7

- (c) 1
- (d) 6

**13. Which of the following is not a periodic property?**

- (a) Metallic character
- (b) Electronegativity
- (c) Oxidizing power
- (d) Hardness

**14. Which among the following is a set of coinage metal?**

- (a) Cu, Ag, Au
- (b) Cu, K, Na
- (c) Au, Pt, W
- (d) Li, Na, B

**15. Which among the following is not radioactive?**

- (a) F
- (b) Th
- (c) Fr
- (d) At

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